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J. Richard Elliott, Carl T. Lira

J. Richard Elliott is Professor of Chemical Engineering at the University of Akron in Ohio. He has taught courses ranging from freshman tools to senior process design as well as thermodynamics at every level. He has worked with the NIST lab in Boulder and ChemStations in Houston. He holds a Ph.D. from Pennsylvania State University.

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J. Richard Elliott is Professor of Chemical Engineering at the University of Akron in Ohio. He has taught courses ranging from freshman tools to senior process design as well as thermodynamics at every level. He has worked with the NIST lab in Boulder and ChemStations in Houston. He holds a Ph.D. from Pennsylvania State University.

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When a solution does not follow the ideal solution approximation we can apply an EOS or the "correction factor", γ_i , yielding the general expression for K-ratio $K_i = \frac{P_i^{\text{sat}}}{P} \frac{\phi_i^{\text{sat}}}{\phi_i} \exp\left[\frac{V_i^{\text{sat}}(P - P_i^{\text{sat}})}{RT}\right]$

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Raoult's Law $f_i^{\text{V}} = f_i^{\text{L}}$ ⇒

Substituting the ideal solution result ⇒

$f_i^{\text{V}} = f_i^{\text{L}} = \phi_i^{\text{L}} \left(\frac{P}{P_i^{\text{sat}}}\right)^{\omega_i}$

for the vapor We correct the fugacity of the saturated liquid for the effect of pressure on the liquid. $\phi_i^{\text{L}} = \phi_i^{\text{L}} \left(\frac{P}{P_i^{\text{sat}}}\right)^{\omega_i}$

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Chapter 14, Interpretation of reaction energy balance from other texts in terms of the notation in Elliott and Lira Chemical Reaction Engineering, 3rd ed., Levenspiel, 1999. Elements of Chemical

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Reaction Engineering, 3rd ed., Fogler, 1999. Introductory Chemical Engineering Thermodynamics Home Page. PDF Viewer.

Supplementary Materials for Students

Chemical Engineering Thermodynamics CHE 3062. All Videos Spring 2020 (this link also contains videos from Polymer Physics class MW lectures at 10:10) M,T,W,R 12:20 to 1:15 Swift 809 (Help Session Wednesdays 3-5 ERC 435) (Nick Patel/Aditya Challa Help Session Wednesdays 6-9pm ERC 405) Professor Greg Beaucage 492 Rhodes Hall beaucag@uc.edu

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solution approximation we can apply an EOS or the "correction factor", γ_i , yielding the general expression for K-ratio

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